

**NANOCOMPOSITE CERAMICS OF OXIDE AND NON-OXIDE PHASES AND
METHODS FOR PRODUCING SAME**

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BACKGROUND OF THE INVENTION

1. Field of the Invention

The present invention involves methods and materials for ceramics and, more
particularly, methods and materials for amorphous nanocomposite ceramics and
10 devices utilizing the same.

2. Statement of the Problem

Formation of ceramics from polymer precursors has received widespread
attention recently, mainly because the processing is done at lower temperatures and
with simpler procedures than conventional processes of sintering ceramic powders.

15 The formation of ceramics from polymer precursors is typically performed by
first thermosetting the polymer precursor to a solid material, and then pyrolyzing the
solid material to form the ceramic.

One of the most interesting groups of such ceramic materials includes
amorphous compounds of silicon, carbon, and nitrogen. Amorphous silicon
20 carbonitride (SiCN) is a relatively new material with potential for a wide range of
applications requiring materials with multifunctional properties. This potential is due to
SiCN being chemically stable at temperatures up to 1500°C, and having excellent
resistance to creep, oxidation, and thermal shock. Bulk form SiCN has been
fabricated from commercially available polymeric precursors. There are significant
25 shortcomings, however, with present SiCN fabrication methods that have limited its
use and prevented many products from being commercially successful. One major
reason for this lack of success lies in the nature of the current process used to make
polymer-based SiCN materials.

The current process for making polymer-based SiCN materials consists of two
30 steps: thermosetting, which is typically polymerization of a liquid form of the precursor
into a rigid plastic body, known as the "green body", followed by pyrolyzing the rigid
plastic into a monolithic SiCN ceramic. These SiCN ceramics are a non-oxide

ceramic, as oxygen has been considered generally to be detrimental to the material.

There are shortcomings with the monolithic SiCN ceramics of the prior art. One is that, being non-oxide, they lack functional properties of existing oxide ceramic monolithic components such as magnets, capacitors, ferroelectric actuators, and others. For this reason, the existing SiCN monolithic ceramics, although having superior mechanical and thermo-mechanical properties over oxide monolithic ceramics, are frequently not suitable replacements.

Another problem with the present process for making polymer-based SiCN materials is the respective temperatures at which thermosetting begins and pyrolysis begins, coupled with the rheology of the known precursors. This substantially limits the scope of shapes, forms, and applications of the SiCN products.

One example of this limitation is apparent from the ongoing quest for SiCN fiber. The search for SiCN fiber is not new, as it has been known, in theory anyway, that such fiber could replace graphite fiber for many applications with increased scope of use and improvement in performance. The reason is that graphite exhibits oxidation, devitrification, and degradation above about 800°C in air. SiCN is stable at considerably higher temperatures. The difficulty, though, is that acceptable quality fibers of SiCN ceramic are difficult to economically produce using the current methods. The fabrication is difficult because, ideally, the precursor would have a rheology suitable for fiber drawing at a temperature just below the thermosetting temperature, and that it have an onset of pyrolysis at a temperature just above the thermosetting temperature. This would allow the fiber to be drawn and immediately thereafter thermoset into a rigid form, which could then be pyrolyzed without losing its shape. Known precursors of SiCN, however, do not have these qualities.

Others have attempted SiCN-type fibers, or alternates, as part of this quest to replace graphite. In the mid-1990s, the Bayer company in Germany announced it had a process for drawing fibers from silicon boron carbonitride (SiBCN); however, this process is known in the art to be too expensive and has not evolved into a successful commercial venture. Nicalon has been used, but Nicalon fibers devitrify at about 1100°C to 1300°C.

SOLUTION

The present invention advances the art and overcomes the aforementioned problems by a composite of nanoscale oxide ceramic phases dispersed in a non-oxide ceramic matrix material. The composite achieves new synergies in the properties of the composite, not only combining the properties of the oxide and non-oxide materials into one composite material, but also providing a new genre of materials where the nanoscale dispersion of the oxide phase leads to novel properties that cannot be obtained in the coarser microstructure of the monolithic oxide materials. In a preferred embodiment, the non-oxide ceramic is silicon-carbon-nitrogen-based and the matrix of this phase imparts resistance to mechanical degradation, resistance to chemical degradation, and resistance to oxidation at temperatures up to 1800°C. The nanodispersed oxide phase imparts other "functional" properties, in addition to the high temperature properties of the matrix, to the composite, including: (a) tailored coefficient of thermal expansion; (b) superparamagnetic properties up to very high temperatures; (c) super ferromagnetic properties; (d) superdielectric properties; and (e) superpiezoelectric and electrostrictive properties, etc. The term "super" is applied to these composites for two reasons: (i) because the composites have mechanical and chemical durability at high temperatures, which cannot be sustained in the monolithic oxide materials; and (ii) because the nanoscale dispersion of the oxide phase often leads to novel functional behavior that is not obtained in microscale, monolithic, polycrystalline oxide materials.

The dispersion of the functional oxide ceramics in an amorphous non-oxide matrix of silicon carbon and nitrogen is readily obtained via the methods of this invention, and provides functional as well as mechanical properties superior to, and additional to, those found in their monolithic counterparts. These nanocomposites can replace functional oxide ceramic monolithic components such as magnets, capacitors, ferroelectric actuators, and others. Industries including mechanical, electrical, electronic, telecommunication, aerospace, and others will find wide applicability of this invention. This invention is, therefore, a paradigm shift for the functional ceramic monoliths.

One embodiment of the invention includes a nanoscale dispersion of predominantly crystalline oxide phases in a predominantly amorphous matrix of a non-

oxide ceramic phase. In the preferred embodiment, the non-oxide ceramic phase is composed primarily of the elements silicon, carbon, and nitrogen, but may contain other dopants, such as boron, in order to control the properties of the matrix phase. The phrase "silicon-carbon-nitrogen based material" is defined herein as a
5 predominantly amorphous matrix of a non-oxide ceramic phase, composed primarily of the elements silicon, carbon, and nitrogen, but which may contain other dopants.

The dispersed oxide phase includes, but is not limited to, zirconia, alumina, spinels (e.g., nickel iron oxides), oxides of iron, perovskites (e.g., barium titanate), ceramics with piezoelectric properties (e.g., PZT), dielectric materials (e.g., barium
10 strontium titanate), other perovskites, sometimes referred to as ABO_3 -type materials, and any other suitable oxide ceramics.

Another embodiment of the invention includes a material containing a nanodispersion of zirconia in a SiCN matrix. The matrix phase imparts resistance to mechanical deformation, resistance to oxidation, and resistance to chemical
15 degradation at temperatures up to 1800°C.

In the preferred embodiment, the polymeric precursor materials include silanes, silazanes, and polysilazanes which result in SiCN ceramics upon crosslinking and pyrolysis. The composition of the ceramic product can be varied by appropriate selection of the polymeric precursor material and the pyrolysis environment, and to a
20 lesser extent, by appropriate selection of the casting conditions. One suitable polysilazane is CerasetTM, manufactured and distributed by Kion Corporation, Columbus, Ohio. In accordance with the present invention, the crosslinking can be accomplished by any suitable polymerization reaction known in the art that yields the desired crosslinked polymeric structure. Pyrolysis under argon or nitrogen at
25 temperatures typically less than approximately 1400°C results in SiC_xN_y , where x and y can be varied using a mixture of ammonia and argon.

Another embodiment of the invention includes a fiber formed of the nanodispersion of zirconia in a SiCN matrix. The zirconia phase allows drawing fibers from a commercial source of the polymer that is used in the fabrication of SiCN, such
30 as CerasetTM. These fibers have a far superior chemical stability at high temperatures as compared to presently available non-oxide fibers, collectively known as Nicalon

fibers.

Still another embodiment of the invention includes a nanodispersion of iron oxide in a SiCN matrix, which exhibits remarkable magnetic properties, and superparamagnetic behavior, normally not seen in monolithic ferromagnetic oxide-based ceramics. One aspect of this embodiment is the polymer-derived SiCN matrix having chemical stability at elevated temperatures and excellent resistance to creep, oxidation, and thermal shock. Ferromagnetic ceramics like Fe_2O_3 and Fe_3O_4 have poor mechanical strength, high coercivity, and high hysteresis loss.

A further aspect of the invention includes a nanocomposite ceramic having a tailored coefficient of thermal expansion. The nanocomposite may be a nanodispersion of zirconia in a SiCN matrix. The coefficient of thermal expansion of SiCN is tailored by the incorporation of 1 weight percent to 99 weight percent zirconium oxide. Yet another aspect of the invention is a sealing material for multilayer fuel cell structures at high temperature, comprising a nanocomposite ceramic having a tailored coefficient of thermal expansion. The nanocomposite may be a nanodispersion of zirconia in a SiCN matrix.

Another aspect of the invention is a nanodispersion of barium-strontium-titanate in a SiCN matrix, which has superdielectric properties as well as superior mechanical structure and thermal stability.

Numerous other features, objects and advantages of the invention will become apparent from the following description when read in conjunction with the accompanying drawings.

BRIEF DESCRIPTION OF THE DRAWINGS

FIG. 1 is a photomicrograph of a nanocomposite Si-C-N-Zr-O fiber specimen obtained via a method of the present invention;

FIG. 2 is a graph comparing the oxidation resistance of a nanocomposite Si-C-N-Zr-O fiber via a method of the present invention with that of a pure SiCN fiber, each obtained via a method of the present invention;

FIG. 3 is a graph comparing the thermal stability of the Si-C-N-Zr-O fiber prepared via a method of the present invention with that of a commercially available Nicalon fiber (Nicalon NL202);

FIG. 4 is a photomicrograph of a Si-C-N-Zr-O fiber specimen obtained via a method of the present invention; and

FIG. 5 is a graph comparing the hysteresis curve of a prior art ferrite magnetic material with that of an amorphous nanocomposite Si-C-N-Fe-O magnetic material specimen obtained via a method of the present invention.

DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENT

FIG. 1 is a photomicrograph of a nanocomposite Si-C-N-Zr-O fiber specimen, comprising a nanodispersion of zirconia in a SiCN matrix, obtained via the following described method of the present invention. The specimen was produced from commercially available CerasetTM and zirconium propoxide (Zr-n-p) dissolved in propanol. They are the sources of SiCN and ZrO₂, respectively. The yield of SiCN and ZrO₂ from CerasetTM and Zr-n-p solution, after pyrolyzing at 1000°C in N₂, was measured to be 78% and 28% weight percentage respectively. Based on this yield, Ceraset and Zr-n-p solution were taken in appropriate proportion so as to get 10% volume percentage of ZrO₂ in the fiber after pyrolysis at 1000°C in N₂. First, the CerasetTM and the Zr-n-p solution were mixed together and heat treated at 160°C to yield viscous liquid. The fiber was then drawn from this viscous liquid at room temperature. The fiber then was thermoset to a rigid solid above 160°C and pyrolyzed at 1000°C.

The SiCN matrix phase of the Si-C-N-Zr-O fiber of the invention imparts resistance to mechanical deformation, resistance to oxidation, and resistance to chemical degradation at temperatures up to 1800°C. The invention's addition of the zirconia phase changed the rheology of the SiCN precursor to enable drawing of fibers from CerasetTM and other commercially available sources of the polymer used in the fabrication of SiCN. Otherwise, the fiber drawing from these commercially available precursors is not possible.

As can be seen from the FIG. 1 photomicrograph, the surface of the specimen Si-C-N-Zr-O fiber is dense and free from defects. X-ray diffraction confirmed both the SiCN and ZrO₂ as being in an amorphous phase. The fracture strength and Young's modulus of Si-C-N-Zr-O fiber were evaluated to be 2.6 GPa and 160 GPa respectively, with a fiber diameter of 11 μm. As a comparison, the highest reported fracture

strength of SiCN fiber from laboratory derived precursor is 2.5 GPa, with a fiber processing involving expensive γ -ray curing.

FIG. 2 is a graph comparing the oxidation resistance of pure SiCN with a nanocomposite Si-C-N-Zr-O fiber of the invention, formed of SiCN with a nanodispersion of 10% ZrO₂ by volume. As seen in the FIG. 2 graph, the Si-C-N-Zr-O fiber has significantly better oxidation resistance than SiCN alone. Accordingly, the Si-C-N-Zr-O fiber formed by the invention, using the readily prepared modified precursor, possesses both excellent and superior fracture strength and oxidation resistance compared to SiCN fiber.

FIG. 3 is a graph comparing the thermal stability of the Si-C-N-Zr-O fiber of the invention with that of a commercially available Nicalon fiber (Nicalon NL202). The experiment was carried out under identical conditions for both cases. Under these identical experimental conditions, initiation of thermal degradation for commercially available Nicalon fiber starts at 1300°C. On the other hand, Si-C-N-Zr-O fiber is stable up to approximately 1500°C. Moreover, the weight loss at 1600°C is only 5% for Si-C-N-Zr-O fiber, while that for Nicalon fiber is observed to be approximately 20%. Thus, the present invention's Si-C-N-Zr-O fibers have a far superior chemical stability at high temperatures as compared to presently available non-oxide fibers such as Nicalon fibers.

FIG. 4 is a photomicrograph of another SiCN – ZrO₂ fiber specimen obtained via the present invention. The specimen shown in FIG. 4 was pyrolyzed at 1300°C. The surface shows uniform dispersion of nanoparticles of ZrO₂ in a SiCN matrix.

The Si-C-N-Zr-O fiber of the present invention is contemplated to have significantly improved performance and a much wider scope of applications compared to the currently used graphite and/or Nicalon fibers. The contemplated applications include those involving extreme environments of temperature and/or chemical reactants, including those causing oxidation. These are important because graphite and Nicalon typically suffer from oxidation, devitrification and degradation in such environments. Graphite oxidizes (burns) above about 800°C in air, while Nicalon fibers degrade by devitrification at about 1100°C to 1300°C. The Si-C-N-Zr-O fibers of the invention are stable at temperatures up to 1500°C in an air environment.

Particular contemplated applications of Si-C-N-Zr-O fibers of the present invention include materials for brakes in aircraft, where the current practice is to use graphite fibers, heat exchangers in energy conversion systems, and applications in space technologies. Another embodiment of the invention is an amorphous nanocomposite Si-C-N-Fe-O soft ferrite magnetic material. FIG. 5 is a graph illustrating one of the benefits of this material. The graph of FIG. 5 plots induced magnetization in Gauss as function of applied field in Oesterds. The inset shows the same graph for a prior art material, Fe_3O_4 , and a clearly drastic improvement in a sample's hysteresis loss when compared to a sample of the prior art ferrite magnetic material.

The FIG. 5 sample composite was made by a polymer derived route using powdered Fe_3O_4 obtained from Fisher Scientific, Fair Lawn, New Jersey, and CerasetTM, obtained from Kion Corporation, Columbus, Ohio. The powdered Fe_3O_4 was dispersed in liquid CerasetTM using an ultrasonic bath. The dispersion was heat treated at 400°C in a nitrogen environment to crosslink the precursor mixture. The heat treated composition was ball milled, followed by pelletization by warm pressing at 350°C and 30 MPa. The pellet was then pyrolyzed under a flowing nitrogen environment at 1000°C, with very slow heating and cooling rates. The mixing ratio of powdered Fe_3O_4 to liquid CerasetTM was such that the final pyrolyzed ceramic composition was Fe_3O_4 – 70% and SiCN – 30% by volume in final composite.

As seen from the FIG. 5 "Magnetization vs. Applied Field" curve, the amorphous nanocomposite SiCN – Fe_3O_4 of this invention has near zero hysteresis. Further, the FIG. 5 curve for the ferrite shows a coercive force of about 1000 Oesterds, while the nanocomposite exhibits a coercive force of only 10 Oesterds.

The nanocomposite of SiCN and ferrite of this invention has remarkable properties which have never before been seen in monolithic ferrites, including: (a) ten to two hundred times the permeability of monolithic polycrystalline ferrites; and (b) nearly zero coercive field and negligible hysteretic loss.

Further, the SiCN – Fe_3O_4 composite can be fabricated by this invention at low temperatures such as, for example, less than 1000°C. In comparison, monolithic ferrites are prepared by the sintering process at much higher temperatures (1200°C to

1400°C). The sintering process often employs sintering aids that can degrade the properties of the material. The polymer derived process of the invention does not involve any sintering aids.

Still further, the polymer-derived SiCN matrix of this embodiment has chemical stability at elevated temperatures and excellent resistance to creep, oxidation, and thermal shock. Ferromagnetic ceramics, like Fe_2O_3 and Fe_3O_4 , have poor mechanical strength. The fracture strength of SiCN – Fe_3O_4 nanocomposites was measured to be 175 MPa. This composite does not exhibit any degradation in magnetic properties when in use at a temperature of approximately 500°C in air. Therefore, the SiCN – Fe_3O_4 nanocomposite of this invention has these benefits in addition to its clearly superior coercivity and hysteresis characteristics.

The soft ferrite nanocomposite SiCN and ferrite materials produced by the methods of this invention are contemplated to have extensive applications including, for example, without deflection yokes of cathode ray tubes (CRT), power switch transformers, retro-sweeping transformer for televisions, radio antennae, chokes, rotary transformers of audio visual (AV) machines, ballast of energy saving lights, and transformers.

A further aspect of the invention is attained by using zirconium oxide as the oxide phase of the oxide/non-oxide nanodispersion ceramic of the invention. Zirconium oxide provides selective tailoring of the coefficient of thermal expansion of the SiCN matrix, ranging from 1 weight percent to 99 weight percent zirconium oxide. A contemplated product of zirconium oxide as the oxide phase nanodispersed in the non-oxide SiCN is a sealing material for multilayer fuel cell structures, usable at high temperatures.

Another aspect of the invention is a nanodispersion of barium-strontium-titanate in a SiCN matrix, which is predicted by the present inventors as likely having superdielectric properties as well as superior mechanical structure and thermal stability.

Particular contemplated applications of the Si-C-N-Zr-O system also include multilayer coating systems in high temperature components such as blades, combustors, nozzles, and linings in gas turbine engines. The polymer route to

processing and the nanoscale microstructure of these coatings can be an advantage in providing thermal and environmental barriers for higher performance in high temperature and aggressive environments.

Each of the above examples shows a different and novel aspect of the composite materials according to the present invention. The scope of this invention, however, is not limited to these examples but extends generally to composites that are constructed from the SiCN-based non-oxide matrix, and the broad range of oxide ceramics described above. The present invention advances the art by dispersing crystalline oxide ceramics at nanometer scale in noncrystalline, non-oxide ceramics to impart various functional properties to the composite. The functional properties exhibited by the composite far exceed those predictable, with any reasonable degree of certainty, by a simple rule of mixtures for composites. These composites, according to the invention, exhibit better mechanical properties than their monolithic counterparts. Further, the invention's methods of dispersing functional oxide ceramics in an amorphous non-oxide matrix are readily carried out.

It should be understood that the particular embodiments shown in the drawings and described within this specification are for purposes of example and should not be construed to limit the invention which will be described in the claims below.